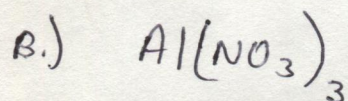
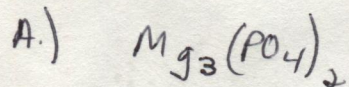


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① Write the solubility product expressions for



② The solubility of $\text{Cr}(\text{OH})_3$ is 4.8×10^{-6} moles/L at 25°C .
What is its K_{sp} at this temperature?

③ At 25°C , a saturated solution of PbCl_2 contains 6.2×10^{-6} mole of compound in one liter of solution. What is the K_{sp} of PbCl_2 ?

④ In a saturated solution at 25°C , what is the molar solubility of MgF_2 ($K_{sp} = 7.8 \times 10^{-8}$)?

(5) If the solubility of AlPO_4 is $2.5 \times 10^{-4} \text{ g/L}$, what is the K_{sp} ?

326 To 01x 1.4 at 40°C and to 0.0017 at 25°C.
The solubility product of AlPO_4 is 1.7×10^{-25} .