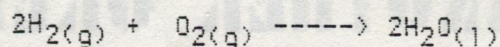


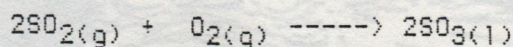
"LETS STOICH AGAIN"



1) TO PRODUCE 0.600 MOLE OF $\text{H}_2\text{O}(\text{l})$:

a) HOW MANY MOLES OF $\text{H}_2(\text{g})$ ARE NEEDED?

b) HOW MANY LITERS OF $\text{O}_2(\text{g})$ ARE NEEDED?



2) HOW MANY MOLES OF $\text{SO}_3(\text{l})$ WILL BE FORMED REACTING WITH 2.1973×10^{20} MOLECULES OF $\text{O}_2(\text{g})$?



3) HOW MANY GRAMS OF BARIUM CHLORIDE WILL BE ~~PRODUCED~~ ^{reactant} ~~FROM~~ 188 GRAMS OF ALUMINUM SULFATE?
needed to produce

4) IF 78 LITERS OF OXYGEN ARE CONSUMED WHEN A CANDLE (PARAFFIN $\text{C}_{25}\text{H}_{52}$) IS BURNED, WHAT VOLUME OF CARBON DIOXIDE IS PRODUCED?

5) WHEN 4 MOLES OF SODIUM PHOSPHITE REACTS WITH 4 MOLES OF ZINC PERBROMATE, HOW MANY MOLES OF SODIUM PERBROMATE ARE PRODUCED?